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Add- 4-E/10, Dabouli-I, Near Durga Mandir

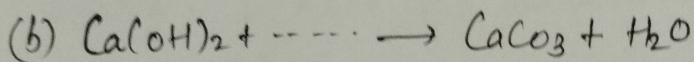
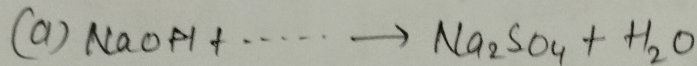
CLASSES - V TO XII (CBSE, ICSE/ISC), IIT/NDA/TGT/PGT

By- SUSHEEL BHATT (MOB- 6306893082) FOUNDER & FACULTY OF MATHEMATICS

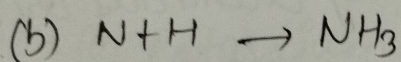
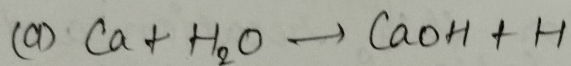
CLASS- Xth

TOPIC → CHEMICAL REACTIONS AND EQUATION

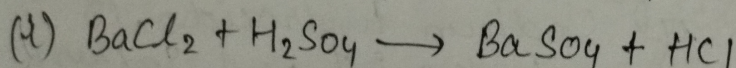
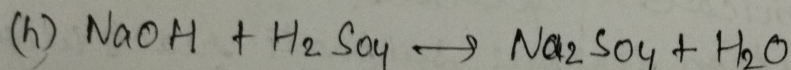
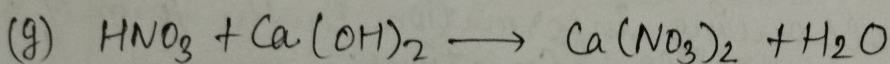
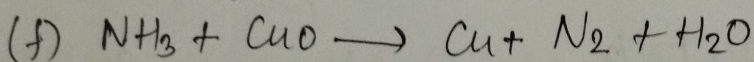
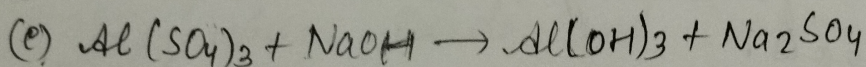
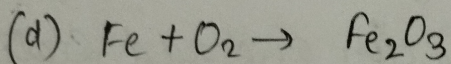
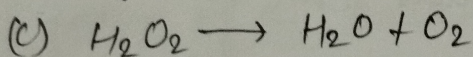
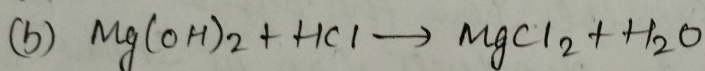
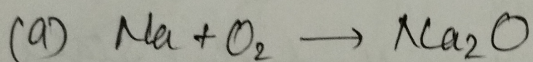
Q1 → Complete and balance the following Equations:-



Q2 → Correct and Balance of the following Equations:



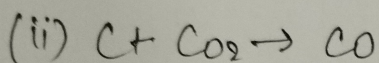
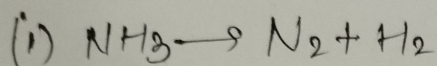
Q3 → Balance the following Equations:-



Q.44 (a) What is a Chemical Equation? Explain with the help of an Example.

(b) Giving Examples, state the difference between Balanced and unbalanced chemical Equations.

(c) Balance the following chemical Equations.



Q.51 When hydrogen is passed over copper Oxide, Copper and Steam are formed. Write a balanced Equation for this reaction and state which of the following chemicals are:

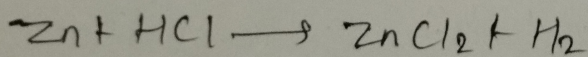
(a) Elements (b) Compounds

(c) reactants (d) products

(e) metal (f) Non-metal

Q.61 (a) Explain, with Example, how the physical States of the reactants and products can be shown in a chemical Equations.

(b) Balance the following Equation and add State Symbols:



(c) Convey the following information in the form of a balanced chemical Equations:

"An aqueous solution of ferrous sulphate reacts with an aqueous solution of sodium hydroxide to form a precipitate of ferrous hydroxide and sodium sulphate remains in solution".

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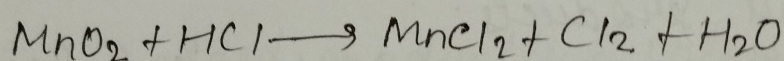
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Q.73 (a) Aluminium hydroxide reacts with Sulphuric acid to form aluminium Sulphate and water. Write a balanced Equation for his reaction.

(b) Balance the following Chemical Equation:



Q.8 (a) Potassium chlorate (KClO_3) on heating forms Potassium chloride and Oxygen. Write a balanced Equation for his reaction and indicate the Evolution of gas.

(b) Rewrite the following information in the form of a balanced chemical Equation:

magnesium burns in Carbon dioxide to form magnesium Oxide and Carbon.

Q.8 (c) Substitute formulae for names and balance the following Equations:

Calcium carbonate reacts with hydrochloric acid to produce calcium chloride, water and Carbon dioxide gas.

(b) Write a balanced chemical Equation with State Symbols for the following reaction.

Sodium hydroxide solution reacts with hydrochloric acid solution to produce sodium chloride solution and water.

Q.9 → Ammonia reacts with Oxygen to form nitrogen and water. Write a balanced Chemical Equation for this reaction. Add the State Symbols for all the reactants and products.

Q.10 → Write a balanced chemical Equation for the process of photosynthesis giving the physical states of all the substances involved and the conditions of the reaction.

Q.11 → (a) What is meant by a Chemical reaction? Explain with the help of an example.

(b) Give one example each of a Chemical reaction characterised by:

- (i) Change of a gas
- (ii) Change in colour
- (iii) Formation of a precipitate
- (iv) Change in temperature
- (v) Change in state

Q.12) (a) State the various characteristics of chemical reactions:

(b) State one characteristic each of the chemical reaction which takes place when:

(i) dilute hydrochloric acid is added to sodium carbonate

(ii) lemon juice is added gradually to potassium permanganate solution.

(iii) dilute sulphuric acid is added to barium chloride solution.

(iv) quicklime is treated with water

(v) wax is burned in the form of a candle.

Q.13) (a) what do you understand by exothermic and endothermic reactions?

(b) Give one example of an exothermic reaction and one of an endothermic reaction.

(c) which of the following are endothermic reactions and which are exothermic reactions?

(i) Burning of natural gas (ii) Photosynthesis

(iii) Electrolysis of water (iv) Respiration.